

Indium nitride crystals with flower-like structure

Naoyuki Takahashi,* Arei Niwa, Haruka Sugiura and Takato Nakamura

*Department of Materials Science and Technology, Faculty of Engineering, Shizuoka University, 3-5-1 Johoku, Hamamatsu 432-8561, Japan. E-mail: tntakah@ipc.shizuoka.ac.jp; Fax: +81-53-478-1197; Tel: +81-53-478-1197**Received (in Cambridge, UK) 12th November 2002, Accepted 23rd December 2002**First published as an Advance Article on the web 14th January 2003*

Preparation of indium nitride at atmospheric pressure has been examined by means of halide chemical vapour deposition; from the SEM observations of the crystals deposited onto an Si(100) substrate it was found that they showed flower-like structure.

Science and technology of three-dimensionally ordered materials have led to great excitement during the last decade. Amongst them, subjects dealing with very tiny objects are of immense academic interest. There has already been much progress in the synthesis, assembly and fabrication of three-dimensionally ordered materials,¹ and also potential applications of such materials to a wide variety of technologies, such as glancing angle deposition (GLAD)² and template selectivity growth.³

Recently, we examined the preparation of InN utilising a reaction of gaseous InCl₃ and NH₃ under atmospheric pressure (AP-HCVD).⁴ It was found that when the substrate temperature was in the range 723–823 K that indium nitride crystals with unusual flower-like structures were formed. Here, we report the results of the investigation on indium nitride with flower-like structure grown on an Si(100) substrate by means of atmospheric pressure halide chemical vapour deposition.

A horizontal hot-wall quartz reactor was used for the preparation of InN. The nitride was deposited onto an Si(100) substrate under atmospheric pressure. The growth temperature was measured just under the substrate stage with an error limit

of 0.2 K. The substrate used was an n-type Si(100) wafer with a resistance of $5 \times 10^{-4} \Omega \text{ m}$. InCl₃ in a source boat was evaporated at a temperature of 523 K, and supplied to the growth zone of the reactor by purified N₂ carrier gas. NH₃ was also supplied to the growth zone of the reactor by purified N₂ carrier gas. The partial pressures of InCl₃ and NH₃ were adjusted independently by varying the flow rates of N₂ carrier gas. The partial pressure of NH₃ was varied in the range 9.3×10^3 – 2.8×10^4 Pa at a constant partial pressure of InCl₃ of 4.7×10^2 Pa. The purity of the N₂ and NH₃ gases (Air Liquide Co. Ltd) are 99.9999% and 99.9995%, respectively. The purity of InCl₃ (Kojundo Chemical Lab. Co., Ltd) is 99.99%. The

Table 1 Typical growth conditions

Substrate	Si(100)
InCl ₃ partial pressure/Pa	4.7×10^2
NH ₃ partial pressure/Pa	2.8×10^4
Carrier gas	N ₂
Total flow rate/m ³ s ⁻¹	1.25×10^{-5}
Growth temperature/K	823

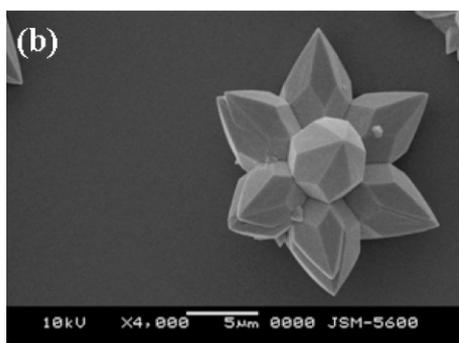
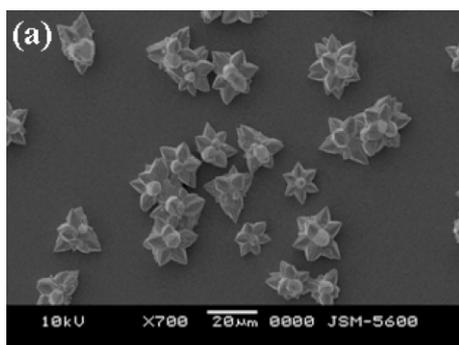


Fig. 1 SEM image of the InN grown under condition of Table 1. (a) magnification $\times 700$, (b) magnification $\times 4000$.

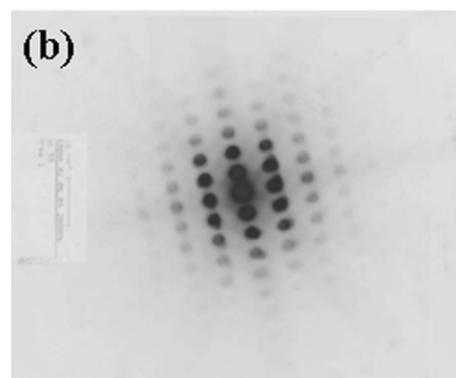
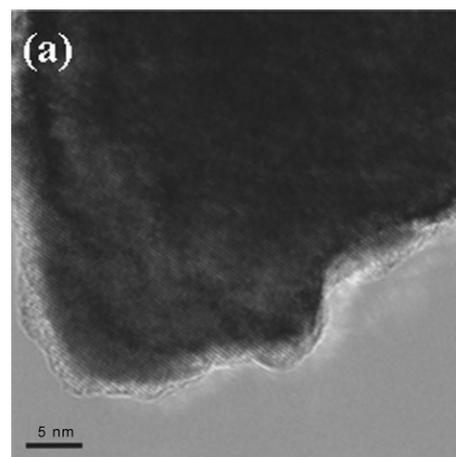


Fig. 2 (a) TEM and (b) electron diffraction images of an InN crystal.

crystallographic structure of the deposited InN was examined by a Rigaku Rint 2000 X-ray diffractometer. Their crystallinity was evaluated by scanning electron microscopy (SEM), transmission electron microscopy (TEM) and electron diffraction.

The crystal structure and orientation of the as-grown films were examined by means of the X-ray diffractometry (XRD). Only the diffraction peaks of InN and the Si substrate were observed. The observed diffraction lines reveal that the hexagonal structure is the only crystalline phase, exhibiting sharp and symmetric reflections for basal (002) and (004) planes, and broad and asymmetric reflections for the non-basal (101), (103) and (110) planes.⁵ The lattice constant was estimated to be 0.570 nm utilising the observed (002) diffraction, which implies that the obtained lattice constant is similar to the reported value of 0.57033 nm for the polycrystalline InN powder.⁵

Fig. 1(a) and (b) show SEM micrographs of the InN grown on the Si(100) substrate under the conditions listed in Table 1. At a glance of the SEM micrographs in Fig. 1(a), it is immediately noticed that the deposited InN has an interesting geometry with flower-like structure. The enlarged SEM micrograph of a flower-like InN arrangement shows that it consists of six petal crystals and a style crystal (Fig. 1(b)). One can see that there is a sixfold axis along the style, and that the respective crystals constituting the petals and a style are staggered hexagonal bipyramids. The flower-like InN was formed only when the substrate temperature was kept at 823 K. As for the partial pressure of NH₃, it was found that the geometry with flower-like structure was hardly affected in the range of $p(\text{NH}_3) = 9.3 \times 10^3$ – 2.8×10^4 Pa at a constant partial pressure of InCl₃ of 4.7×10^2 Pa examined in this study.

A representative TEM micrograph and an electron diffraction pattern of the indium nitride crystals are shown in Fig. 2(a) and (b), respectively, in which the top of the one of the staggered hexagonal bipyramidal crystals was measured. In Fig. 2(a), it is seen that there are a number of lines with a constant spacing of 0.2853 nm, which is assigned to the 002 spacing based on both the JCPDS data⁵ and the XRD results mentioned above. Also, an intense spot pattern appearing in the electron diffraction image is a characteristic pattern of hexagonal structures (Fig. 2(b)). This means that the staggered hexagonal bipyramidal crystals are single crystals of InN.

In summary, we have found that InN with flower-like structure was deposited onto an Si(100) substrate by AP-HCVD upon reaction of InCl₃ and NH₃. Each crystal constituting the flower is a single crystal with staggered hexagonal bipyramidal structure. The formation of such highly unusual crystals is sensitive to the substrate temperature.

This work was supported by the Japan Society for the Promotion of Science through a Grants-in-aid for Scientific Research (A) No. 13305047.

Notes and references

- 1 P. V. Braun, *223rd ACS National Meeting Orlando, FL, USA, April, 2002*, POLY-254(2002); P. H. Holloway and S. L. Jones, *J. Surf. Anal.*, 1998, **4**, 226.
- 2 M. Seto, K. Westra and M. Brett, *J. Mater. Chem.*, 2002, **12**, 2348.
- 3 B. Q. Wei, R. Vajtai, Y. Jung, J. Ward, R. Zhang, G. Ramanath and P. M. Ajayan, *Nature*, 2002, **416**, 495.
- 4 N. Takahashi, A. Niwa, T. Takahashi, T. Nakamura, M. Yoshioka and Y. Momose, *J. Mater. Chem.*, 2002, **12**, 1573.
- 5 JCPDS File 1999, No. 50-1239.